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The atomic number tells you the number  
of \_\_\_\_\_ in one atom of an element. It  
also tells you the number of \_\_\_\_\_ in a

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neutral atom of that element. The atomic number gives the “identity” of an element as well as its location on the periodic table.

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Group 1 elements in the periodic table are known as the alkali metals. They include lithium, sodium and potassium, which all react vigorously with air and water. Group 2 metals are known as the alkaline earth metals and include beryllium, barium, magnesium, calcium

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After filling the first five rows, we still have  $80 - 54 = 26$  more electrons to accommodate. According to Figure 1.3.1, we need to fill the 6s (2 electrons), 4f (14 electrons), and 5d (10 electrons) orbitals. The result is mercury's electron configuration:  $1s^2 2s^2 2p^6 3s^2 3p^6$

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$4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} = \text{Hg} = [\text{Xe} \dots$

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It has two valence electrons but is in  
Group 1 8. Period 2, group 1 7 element I  
t ha s more tha n 5 0 bu les 8 electron

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and i Group 17. This element is a transition metal that is liquid at room temperature. It has the lowest atomic number of any Group 7 elements. Period 4, alkali metal. It has a higher atomic number than carbon and lower P.

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