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CHAPTER 11: STOICHIOMETRY. UNIT 4:
Chemical Reactions, The Mole,
Stoichiometry and Thermodynamics.
Part B:Stoichiometry. Big Picture Ideas:
The identity of the reactants helps
scientists to predict the products in a

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chemical reaction. Quantitative relationships exist with all chemical reactions that allow scientists to predict amounts of products formed, reactants consumed, and percent yield based on theoretical maximum.

CHAPTER 11: STOICHIOMETRY

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-Stoichiometry (Rough outline of the chapter, please use the book, notes & homework to study.) 11.1 Defining Stoichiometry Vocab • stoichiometry • mole ratio Concepts Using Balanced Equations • Number of Atoms • Number of Molecules • Number of Moles • Mass • Law of Conservation of Mass • Volume

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TEACHER GUIDE AND ANSWERS Study
Guide - Chapter 11 - Stoichiometry

Section 11.1 What is stoichiometry? 1.
true 2. true 3. false 4. true 5. true 6. 2,
2, 64.10 7. 3, 3, 96.00 8. 2, 2, 88.02 9. 4,
4, 72.08 10. methanol and oxygen gas
11. carbon dioxide and water 12. 160.10

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g 13. 160.10 g 14. They are equal. 15. A mole ratio is a ratio between the numbers of moles

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Stoichiometry (Rough outline of the chapter, please use the book, notes & homework to study.) 11.1 Defining

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mole ratio Concepts Using Balanced
Equations • Number of Atoms • Number
of Molecules • Number of Moles • Mass o
Law of Conservation of Mass • Volume
11.2 Stoichiometric Calculations
Concepts Mole-Mole ...

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Stoichiometry | pdf Book Manual ...

368 Chapter 11 • Stoichiometry Section
111.11.1 Objectives Describe the types
of relationships indicated by a balanced
chemical equation. State the mole ratios
from a balanced chemical equation.
Review Vocabulary reactant: the starting
substance in a chemical reaction New
Vocabulary stoichiometry mole ratio

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Chapter 11: Stoichiometry

Chapter 11 Stoichiometry.

stoichiometry. mole ratio. excess reactant. limiting reactant. The study of quantitative relationships between the amounts of.... In a balanced equation, the ratio between the numbers of

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moles.... A reactant that remains after a chemical reaction stops.

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Chemistry Chapter 11 Stoichiometry.
stoichiometry. mole ratio. limiting
reactant. excess reactant. the study of
quantitative relationships between the

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amounts of.... in a balanced equation,
the ratio between the number of
moles.... a reactant that is totally
consumed during a chemical reaction....

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Chapter 11 (Stoichiometry)

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CHAPTER STUDY GUIDE Date Class
Stoichiometry Section 11.1 What is stoichiometry? In your textbook, read about stoichiometry and the balanced equation. For each statement below, write true or false. 2. 3. 4. 5. The study of the quantitative relationships between the amounts of reactants used and the

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amounts of products formed by a chemical reaction

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reactant. limiting reactant. The study of quantitative relationships between the amounts of.... In a balanced equation, the ratio between the numbers of moles.... A reactant that remains after a chemical reaction stops.

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CHAPTER Section 11.1 continued In your textbook, read about mole ratios.

Answer the questions about the following chemical reaction. sodium + iron(III) oxide \rightarrow sodium oxide + iron

$6\text{Na}(s) + \text{Fe}_2\text{O}_3(s) \rightarrow 3\text{Na}_2\text{O}(s) + 2\text{Fe}(s)$ 15. What is a mole ratio? 16. How IS a mole ratio

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written? CA S Q C CYA 17. Predict the number of mole ratios for this reaction. Class 18.

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In this video, we will apply stoichiometric concepts to balanced equations in order to determine mass relationships within

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reactions.

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