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Chapter 11 and 12 Study Guide Molarity Molality moles of solute L of solution moles of solute kg of solvent Percent mass grams of solute x 100 grams of solution Mole Fraction mol solute total moles Density of Solution grams solution mL solution X Solution Formation Electrolytic ions present in

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solution Hydrated ions. interactions keep the ions from recrystallizing Not everything that dissolves in water is an electrolytic solution. o Sugar dissolves.

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Name Date CHAPTER Section 11.1 continued In your textbook, read about mole ratios. Answer the questions about the following chemical reaction. sodium + iron(III) oxide \rightarrow sodium oxide + iron $6\text{Na}(s) + \text{Fe}_2\text{O}_3(s) \rightarrow 3\text{Na}_2\text{O}(s) + 2\text{Fe}(s)$ 15. What is a mole ratio? 16.

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This section shows how the information learned in Chapter 11 can be combined with a model for covalent bonding called the valence-bond model to explain the common bonding patterns of the nonmetallic atoms. The most common of these bonding patterns were listed in Chapter

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Chapter 12 Molecular Structure - An Introduction to Chemistry

School: University of Houston Department: Chemistry Course: Fundamentals of Chemistry 2
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