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Chapter 12 Stoichiometry 127 SECTION 12.1 THE ARITHMETIC OF EQUATIONS (pages 353-358) This section explains how to calculate the amount of reactants required or product formed in a nonchemical process. It teaches you how to interpret chemical equations in terms of interacting moles, representative particles, masses, and gas volume at STP.

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Chapter 12 Multiple Choice Identify the letter of the choice that best completes the statement or answers the question. \_\_\_\_ 1. The first step in most stoichiometry problems is to \_\_\_\_\_. a. add the coefficients of the reagents c. convert given quantities to volumes b. convert given quantities to moles d. convert given quantities to masses \_\_\_\_ 2.

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In Example 12.2.1 and Example 12.2.2, the identity of the limiting reactant has been apparent:  $[\text{Au}(\text{CN})_2]^-$ ,  $\text{LaCl}_3$ , ethanol, and para-nitrophenol. When the limiting reactant is not apparent, we can determine which reactant is limiting by comparing the molar amounts of the reactants with their coefficients in the balanced chemical equation ...

### Chapter 12.2: Stoichiometry of Reactions in Solution ...

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