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Chapter 16 Review Acid Base

Chapter 16 - Acid-Base
Equilibria. 16.1 Acids &
Bases: A Brief Review. -
Arrhenius acids and
bases: -- acid: an H⁺
donor HA H A(aq) (aq)
(aq) -- base: an OH⁻
donor MOH M OH(aq)
(aq) (aq) -

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Brønsted-Lowry acids and bases: -- acid: an H^+ donor $HA \rightleftharpoons H^+(aq) + A^-(aq)$ -- base: an H^+ acceptor $B + H^+(aq) \rightleftharpoons BH^+(aq)$.

Chapter 16 Acid-Base Equilibria - Directory

16.2 Water and the pH Scale. It is crucial to understand the acid-base nature of water itself. Amphiprotic substance: A substance that can act as either

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an acid or a base.

Water is a primary
example of this.

Autoionization

Reaction: -A reaction
that results in the
formation of a cation or
an anion.

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**Chapter 16 - Acid-
Base Equilibria**

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WORKSHEET~ 1. What

two substances are

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always produced by a neutralization reaction?

- a. acid and a base
- b. water and a base
- c. water and an acid
- d. water and a salt

Chapter 16: Leader: Acid and Base Review

Chapter 16. 16.1 Acids and Bases: A Brief Review; 16.2 Bronsted-Lowry Acids and Bases; 16.3 The Autoionization of Water; 16.4 The pH Scale; 16.5 Strong

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Acids and Bases; 16.6
Weak Acids; 16.7 Weak
Bases; 16.8

Relationship Between K_a and K_b ; 16.9 Acid-
Base Properties of Salt
Solutions; 16.10 Acid-
Base Behavior and
Chemical Structure;
16.11 ...

16.1 Acids and Bases: A Brief Review - Chapter 1 | Dr. Fus

Chapter 16 Acids and
Bases 1. Acids were

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recognized primarily from their sour taste. Bases were recognized from their bitter taste and slippery feel on skin.

Chapter 16 Acids and Bases

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16: Acids and Bases
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Ka Kb Kw pH pOH pKa

pKb H+ OH-

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Chapter 16 (Acid-Base Equilibria) - Part 3

Major topics: Arrhenius vs. Bronsted-Lowry definition of acids and bases, conjugate acid/base, acid dissociation constant (K_a), & strong vs weak acids.

Chapter 16 (Acid-Base Equilibria) -

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16.2.3 Related
Strengths of Acids and

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Bases. the stronger the acid, the weaker the conjugate base. the stronger the base, the weaker the conjugate acid. equilibrium favors transfer of proton from stronger acid to stronger base.

16.S: Acid-Base Equilibria

(Summary) -

Chemistry

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CHAPTER 14 REVIEW

Acids and Bases

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SECTION 1 SHORT

ANSWER Answer the following questions in the space provided. 1. Name the following compounds as acids:
sulfuric acid a. H_2SO_4
sulfurous acid b. H_2SO_3
3 hydrosulfuric acid c. H_2S
perchloric acid d. HClO_4
hydrocyanic acid e. hydrogen cyanide 2. H

14 Acids and Bases

Section 16.1 - ACIDS

AND BASES; A BRIEF

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REVIEW • Acids and bases were first recognized by the properties of their aqueous solutions. o For example, acids turn litmus red, whereas bases turn litmus blue.

- Arrhenius recognized that properties of acidic solutions are due to H^+ (aq) ions and those of basic solutions are due to OH^-

chapter 16 - Section 16.1 ACIDS AND

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BASES A BRIEF

REVIEW ...

Ka Kb Kw pH pOH pKa
pKb H+ OH-

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Chapter 16 (Acid- Base Equilibria) - Part 2

Chapter 16: Acid-Base
Equilibria and Solubility

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Equilibria 7. You are asked to go into the lab and prepare an acetic acid - sodium acetate buffer solution with a pH of 4.00 ± 0.02 . What molar ratio of CH_3COOH to CH_3COONa should be used? Subscribe to Course Hero to unlock this document

**Chapter
16-Worksheet with
answers - Chapter
16 Acid-Base ...**

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