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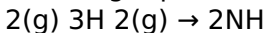
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3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is

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25 g. Calculate the
percentage yield. 2.

6.0 mol of N₂ are
mixed with 12.0 mol of

H₂ according to the
following equation: N



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the left. Show all your work in the space provided. 1. $\frac{\quad}{\quad}$ The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H

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ANSWER Answer the
following questions in
the space provided. 1.
_____ The coefficients
in a chemical equation

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represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products. (c) number of atoms of each element in each compound in a reaction.

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3 PROBLEMS Write the
answer on the line to

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the left. Show all your work in the space provided.

1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield.

2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the following

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Reaction stoichiometry uses molar relationships to determine the amounts of unknown reactants or products from the amounts of known reactants or products.

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Stoichiometry.

SECTION 9.2.

PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many grams of O_2 form if 3.0 mol of KClO_3 are totally

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consumed? 2. Given
the following equation

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the line to the left.

Show all your work in
the space provided. 1.

The following equation
represents a laboratory
preparation for oxygen
gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$

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reaction stoichiometry:
using a balanced
chemical equation to
calculate the number
of grams, moles, or
particles of
reactants/products
involved in a chemical
reaction. Students had
an introduction to
composition
stoichiometry in
Chapter 3 and will now
move on to some more
difficult problems.

Stoichiometry

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Stoichiometry. The Reactant and Product Relationship can be used for prediction if the balanced equation is known. When hydrogen and oxygen combine, $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$ Modern Chemistry Chapter 9 Stoichiometry - Licking $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

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mass, mass to mole,
and mass to Molemov
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of the mole ratios from

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