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first gas, P 2 is the
partial pressure of the

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second gas, and so on.

The kinetic-molecular theory of matter can explain Dalton's law.

CHAPTER 11 Gases

Gases SECTION 3

SHORT ANSWER

Answer the following questions in the space provided. 1. The molar mass of a gas at STP is the density of that gas

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Gases CHAPTER 11
REVIEW Gases

SECTION 3 SHORT

ANSWER Answer the following questions in the space provided. 1.

_____ The molar mass of a gas at STP is the density of that gas (a) multiplied by the mass of 1 mol. (c) multiplied by 22.4 L. (b) divided

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by the mass of 1 mol.

(d) divided by 22.4 L.

2. _____ For the
expression , P

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SECTION 4 Date

CHAPTER 11 REVIEW

Gases Class SHORT

ANSWER Answer the
following questions in
the space provided.

Increasing rate of
effusion. List the
following gases in

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order (Assume all gases are at the same temperature and pressure.) (c) HCl (d) 2. Explain your reasoning for the order of gases you chose in item 1 above. Refer to the kinetic-

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Read the Chemistry in Action section on page 368 of the textbook. Answer questions #1 & #2 at the end of the

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Chemistry in Action
Answer Key

section on page 368.
The Gas Laws Boyle's
Law: the pressure-
volume relationship of
gases. Boyle's Law. ...
Modern Chemistry
Chapter 11 GASES

Modern Chemistry
Chapter 11 GASES
CHAPTER 10 REVIEW
States of Matter
SECTION 3 SHORT
ANSWER Answer the
following questions in
the space provided. 1.

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Match description on the right to the correct crystal type on the left.

b ionic crystal (a) has mobile electrons in the crystal
c covalent molecular crystal (b) is hard, brittle, and nonconducting
a metallic crystal (c) typically has the lowest melting point of the four

**10 States of Matter -
Ms. Agostine's
Chemistry Page**

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Modern Chemistry 74
Quiz Section Quiz: The
Gas Laws In the space
provided, write the
letter of the term or
phrase that best
completes ... 11 Gases
Section: Gases and
Pressure 1. a 2. c 3. d
4. a 5. b 6. c 7. a 8. c 9.
c 10. Section: The Gas
Laws 1. c 2. c 3. b 4. c
5. d 6. 7. d 8. c 9. b 10.
c

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How can you best increase the solubility of a gas in a liquid? a. Increase both the temperature and the pressure. ... Section Quiz: Concentration of Solutions In the space provided, write the letter of the term or phrase that best completes ... Modern Chemistry 84 . Name Section Quiz, continued Class Date 5. In some instances, the

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concentration ...
Answer Key

**Home - David
Brearley High School**

2) At constant volume and temperature, the total pressure exerted by a mixture of gases is equal to the sum of the partial pressures of the component gases
boyle's law 1) The relationship between the pressure and volume of a gas at constant temperature; when volume increase,

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pressure decreases.,
Answer Key

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Characteristics of ...**

Gases spread out and mix with one another, even without being stirred. If the stopper is removed from a container of ammonia in a room, ammonia gas will mix uniformly with the air and spread

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throughout the room.
The random and
continuous motion of
the ammonia
molecules (assump-
tion 3) carries them
throughout the
available space. Such
spontaneous

CHAPTER 10 States Matter

A 500.0 mL gas sample
at STP is compressed
to a volume of 300.0
mL and the
temperature is

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increased to 35.0°C .

What is the new pressure of the gas in pascals? 8. A sample of gas occupies 1000. mL at standard pressure.

What volume will the gas occupy at a pressure of 600. mm Hg if the temperature remains constant? 96

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Answers States Of ...**

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Three of the primary components of air are carbon dioxide, nitrogen, and oxygen. In a sample containing a mixture of only these gases at exactly 1 atm, the partial pressures of carbon dioxide and nitrogen are given as $P_{\text{CO}_2} = 0.285$ torr and $P_{\text{N}_2} = 593.525$ torr. What is the partial pressure of oxygen?

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3 at 25°C is 10. g of solute per 100. g of H₂O. a. If 15 g of KClO₃ are stirred into 100 g of water at 25°C, how much of the KClO₃ will dissolve? Is the solution saturated, unsaturated, or supersaturated? 10 g of KClO₃ will dissolve, but 5 g will not, despite thorough stirring. The solution is saturated.

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SOLUTIONS 105

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Other Results for
Modern Chemistry
Chapter 1 Review
Answers Matter And
Change: ... CHAPTER 1
REVIEW Matter and
Change SECTION 3
SHORT ANSWER

Answer the following
questions in the space
provided. 1. A
horizontal row of
elements in the
periodic table is called
a(n) period. 2. The

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symbol for the element
in Period 2, Group 13,
is B .

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Chapter 1 Review
Answers Matter And
Change**

The study of the
chemical behavior of
gases was part of the
basis of perhaps the
most fundamental
chemical revolution in
history. French
nobleman Antoine
Lavoisier, widely

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regarded as the “father of modern chemistry,” changed chemistry from a qualitative to a quantitative science through his work with gases. He discovered the law of conservation of matter, discovered the role of oxygen in ...

8.3 Stoichiometry of Gaseous Substances, Mixtures, and ...

CHAPTER 1 REVIEW

Matter and Change

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MIXED REVIEW SHORT

ANSWER Answer the following questions in the space provided. 1.

Classify each of the following as a homogeneous or heterogeneous substance. a. sugar d. plastic wrap b. iron filings e. cement sidewalk c. granola bar

2. For each type of investigation, select the most appropriate branch of chemistry from the following

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